

Paper Reference(s) 4CH1/2C

Pearson Edexcel International GCSE (9–1)

Chemistry

Unit: 4CH1

Paper 2C

Wednesday 12 June 2019 – Morning

Time: 1 hour 15 minutes plus your additional time allowance

INSTRUCTIONS TO CANDIDATES

Write your centre number, candidate number, surname, other names and your signature in the boxes below. Check that you have the correct question paper.

Centre No.					
Candidate No.					
Surname					
Other names					
Signature					
Paper Reference	4	C	H	1	/ 2 C



- Use **BLACK** ink or ball-point pen.
- Answer **ALL** questions.
- Answer the questions in the spaces provided – there may be more space than you need.
- Some questions must be answered with a cross in a box ☐. If you change your mind about an answer, put a line through the box ☒ and then mark your new answer with a cross ☐.

MATERIALS REQUIRED FOR EXAMINATION

Calculator, ruler

ITEMS INCLUDED WITH QUESTION PAPERS

Periodic Table

INFORMATION FOR CANDIDATES

- The total mark for this paper is 70.
- The marks for **EACH** question are shown in brackets – use this as a guide as to how much time to spend on each question.

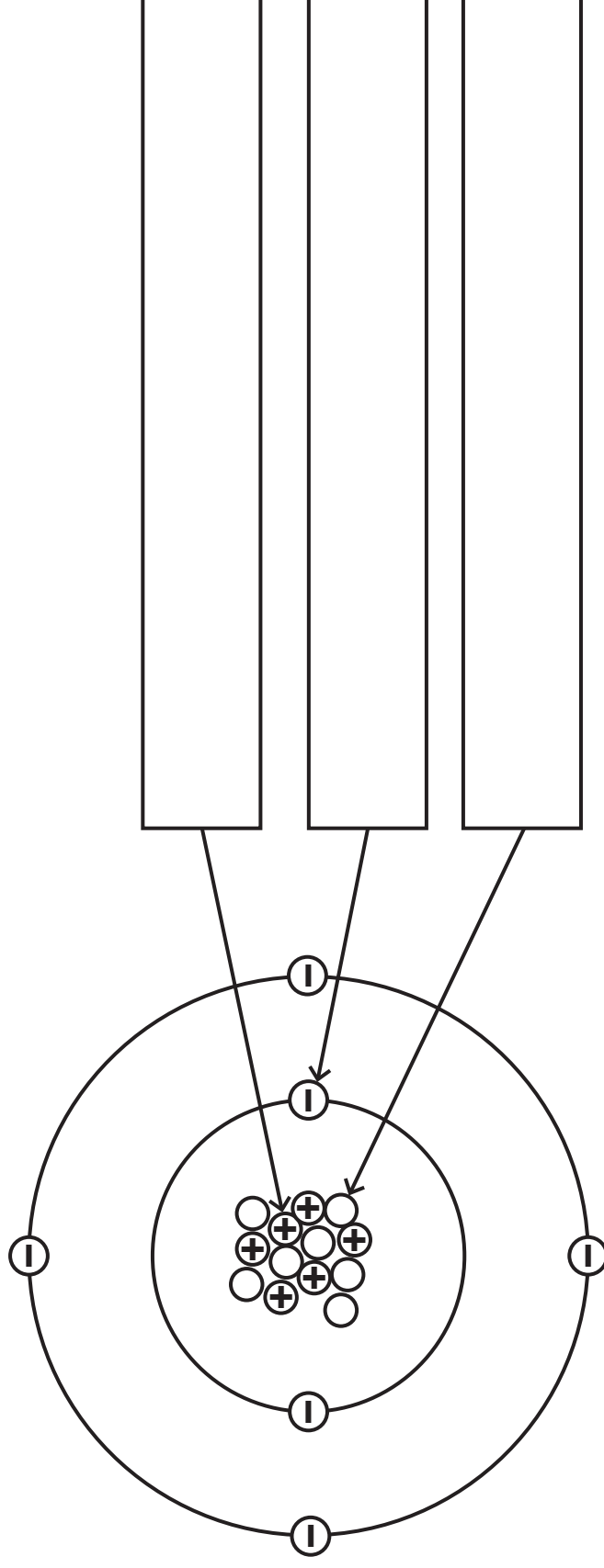
ADVICE TO CANDIDATES

- Read each question carefully before you start to answer it.
- Try to answer every question.
- Check your answers if you have time at the end.

(Turn over)

Answer ALL questions. Write your answers in the spaces provided.

- 1 The diagram shows the particles in an atom of an element.**



- (a) The box gives the names of some particles.**

electron	ion	molecule	neutron	proton
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Use words from the box to label the diagram. (3 marks)

(Question continues on next page)

(Turn over)

(b) Give the mass number of this atom. (1 mark)

(c) Complete the sentence about isotopes. (2 marks)

Isotopes are atoms that have the same number of

but have a different number of

_____ .

(TOTAL FOR QUESTION 1 = 6 MARKS)

(Questions continue on next page)

- 2 The table gives some information about the halogens, chlorine, bromine and iodine.

Halogen	Physical state at room temperature	Colour
chlorine	gas	pale green
bromine		red-brown
iodine	solid	

- (a) Complete the table. (2 marks)
- (b) Chlorine has two isotopes of mass numbers 35 and 37

The relative percentage of each isotope in a sample of chlorine is

chlorine-35 77.78% chlorine-37 22.22%

Calculate the relative atomic mass of this sample of chlorine.

Give your answer to one decimal place. (3 marks)

relative atomic mass = _____

(Question continues on next page)

(Turn over)

- (c) A student is given an aqueous solution of chlorine and an aqueous solution of potassium bromide.**

Explain how he can use these two solutions to compare the reactivity of chlorine with the reactivity of bromine. (4 marks)

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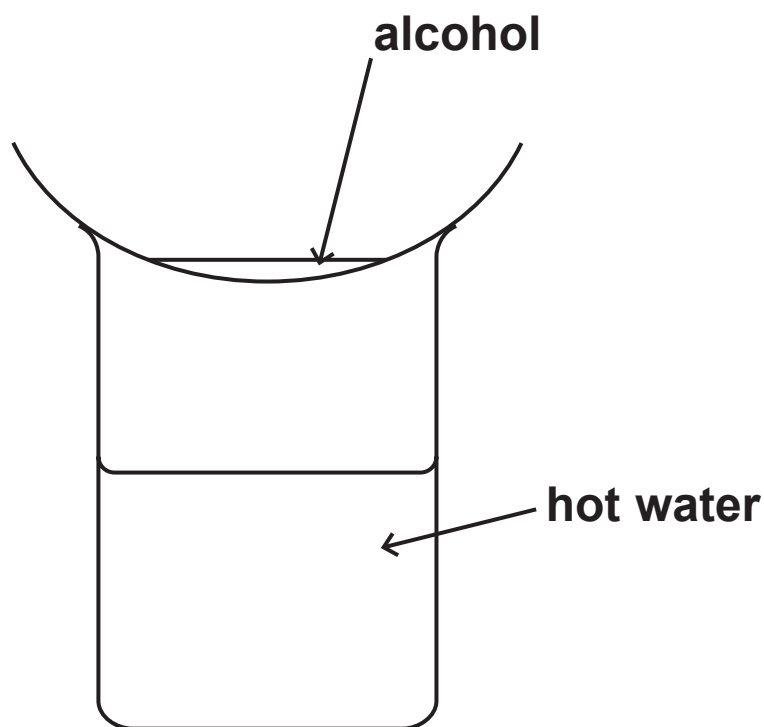
(TOTAL FOR QUESTION 2 = 9 MARKS)

(Questions continue on next page)

(Turn over)

- 3 Methanol, ethanol, propanol and butanol are alcohols. They are all liquids that evaporate easily when warmed.

A student uses this apparatus to compare the time taken for the four liquids to evaporate.



She uses this method.

- pour some methanol into an evaporating basin
- place the evaporating basin on top of a beaker containing hot water
- measure the time taken for the methanol to evaporate completely
- repeat the experiment with each of the other alcohols, using the same apparatus

(Question continues on next page)

(Turn over)

- (a) State two variables the student should control to make sure her results are valid. (2 marks)**

1 _____

2 _____

- (b) State why it is not safe to heat the evaporating basin directly with a Bunsen flame. (1 mark)**

(Question continues on next page)

(Turn over)

- (c) The table shows the results of experiments done by four students, A, B, C and D.

Alcohol	Formula of alcohol	Time taken for liquid to evaporate in s				Mean time in s
		Student A	Student B	Student C	Student D	
methanol	CH ₃ OH	20	24	22	26	23
ethanol	C ₂ H ₅ OH	32	34	35	30	33
propanol	C ₃ H ₇ OH	45	47	50	48	48
butanol	C ₄ H ₉ OH	64	63	90	60	

- (i) Calculate the mean (average) time for butanol to evaporate. (2 marks)

mean time = _____ s

(Question continues on next page)

(Turn over)

- (ii) Explain how the results show which alcohol evaporates most easily. (2 marks)

(Question continues on next page)

- (iii) State the relationship between the number of carbon atoms in the molecule and how easily the alcohol evaporates. (2 marks)

(TOTAL FOR QUESTION 3 = 9 MARKS)

(Questions continue on next page)

(Turn over)

4 This question is about metals.

**(a) Which statement describes metallic bonding?
(1 mark)**

- ☐ **A electrostatic attraction between oppositely charged ions**
- ☐ **B electrostatic attraction between the nuclei of two atoms and a pair of electrons shared between them**
- ☐ **C electrostatic attraction between positively charged particles and delocalised electrons**
- ☐ **D electrostatic attraction between atoms**

(b) Aluminium is malleable and can be easily shaped to make saucepans used for cooking food.

**State two other properties of aluminium that make it suitable for saucepans used for cooking food.
(2 marks)**

1 _____

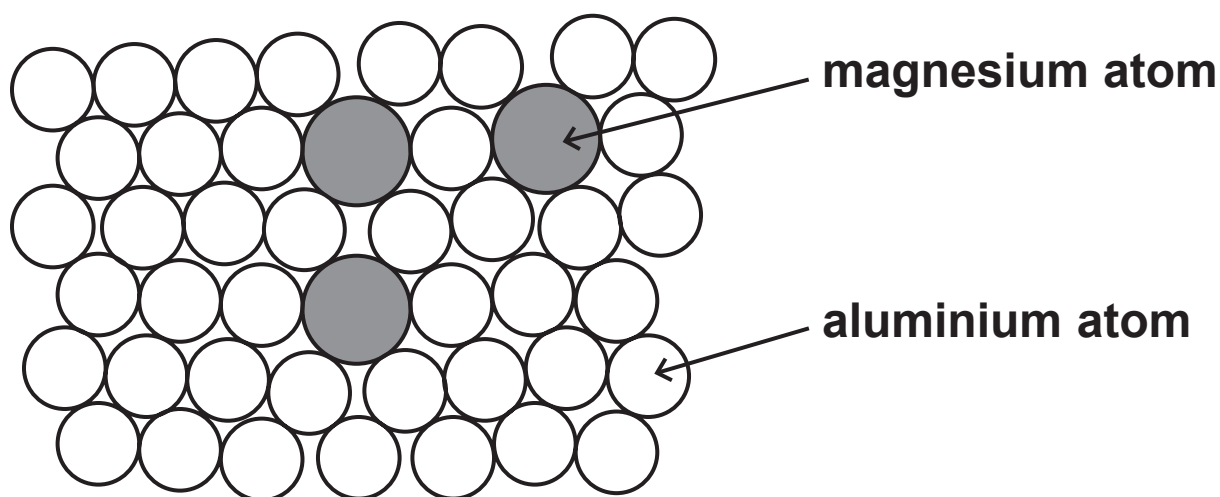
2 _____

(Question continues on next page)

(Turn over)

- (c) **Magnalium is an alloy of aluminium and magnesium.**

The diagram shows how the atoms are arranged in this alloy.



- (i) **State what is meant by the term ALLOY.**
(1 mark)

(Question continues on next page)

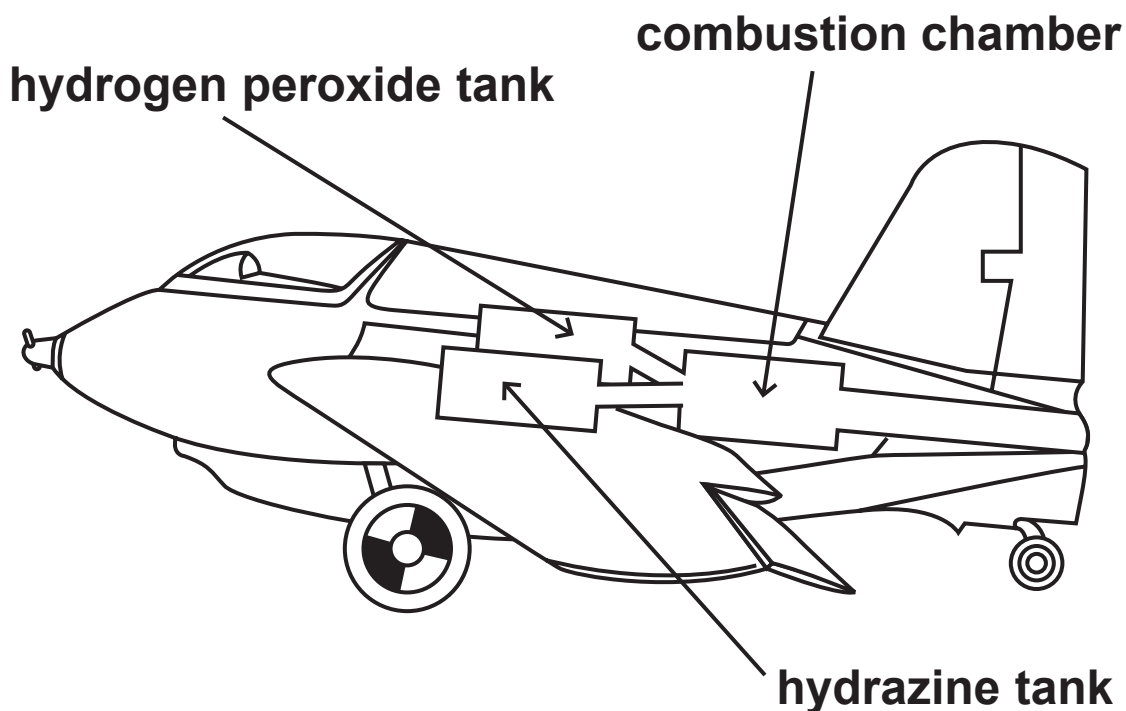
(ii) Explain why magnalium is harder than aluminium. (3 marks)

(TOTAL FOR QUESTION 4 = 7 MARKS)

(Questions continue on next page)

(Turn over)

- 5 During the Second World War, engineers developed a rocket-powered aircraft.



The aircraft carried these two liquids

- hydrazine, N_2H_4
- hydrogen peroxide, H_2O_2

When these two liquids mix in the combustion chamber, they evaporate and then react rapidly to form nitrogen gas, N_2 , and steam, H_2O

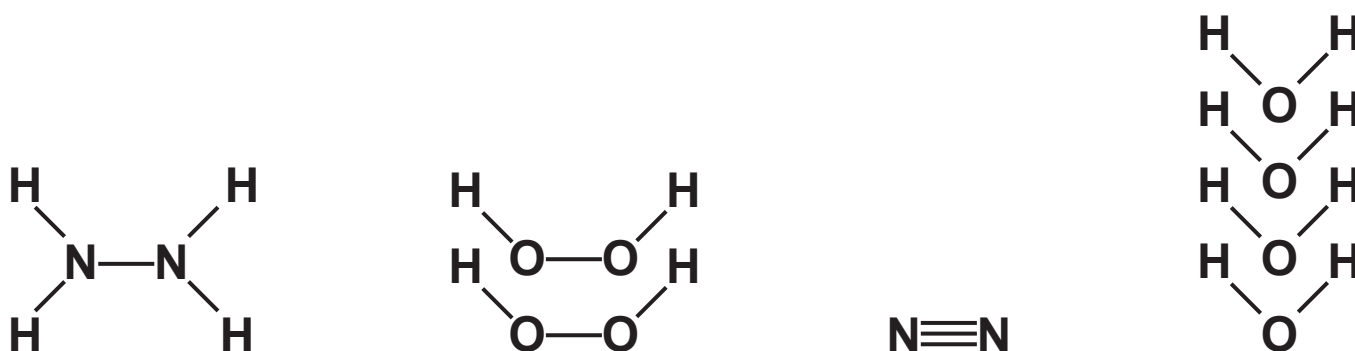
The reaction is exothermic.

(Question continues on next page)

The equation for the reaction is



The displayed formulae for the reactants and products are



- (a) The tables give the bond energies for the bonds broken in the reactants and the bonds made in the products.

Bonds broken	
bond	bond energy in kJ/mol
N—N	159
N—H	391
O—O	143
O—H	463

Bonds made	
bond	bond energy in kJ/mol
N≡N	945
O—H	463

(Question continues on next page)

(Turn over)

- (i) Use the data in the tables to calculate the total amount of energy required to break all of the bonds in the reactants. (1 mark)

energy required = _____ kJ

- (ii) Use the data in the tables to calculate the total amount of energy released when all of the bonds in the products are made. (1 mark)

energy released = _____ kJ

(Question continues on next page)

(Turn over)

(iii) Calculate the enthalpy change, ΔH , in kJ/mol, for the reaction.

Include a sign in your answer. (3 marks)

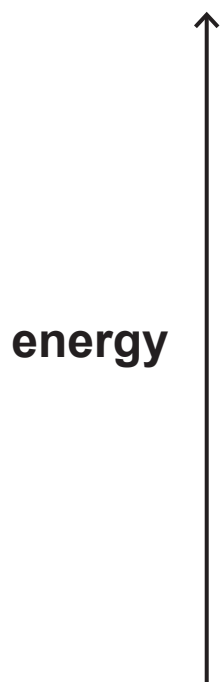
$\Delta H =$ _____ kJ/mol

(Question continues on next page)

(b) Explain, in terms of bonds broken and bonds made, why this reaction is exothermic. (2 marks)

(Question continues on next page)

- (c) Draw an energy level diagram for the reaction between N_2H_4 and H_2O_2 (3 marks)



(TOTAL FOR QUESTION 5 = 10 MARKS)

(Questions continue on next page)

(Turn over)

- 6 Some cars in Brazil use ethanol, $\text{C}_2\text{H}_5\text{OH}$, as a fuel instead of petrol.

The ethanol is made by the fermentation of glucose which is obtained from sugar cane.

The sugar is extracted from the sugar cane and then dissolved in water to make a sugar solution.

- (a) (i) Name the substance that is added to the sugar solution that causes glucose to ferment.
(1 mark)
-
-

- (ii) Which temperature is the most suitable for fermentation? (1 mark)

☐ A 0°C

☐ B 10°C

☐ C 30°C

☐ D 80°C

(Question continues on next page)

(Turn over)

(iii) Explain why fermentation is done in the absence of air. (2 marks)

**(b) (i) State what is meant by the term FUEL.
(1 mark)**

(ii) Write a chemical equation for the complete combustion of ethanol in air. (2 marks)

- (c) Ethanol is also manufactured by reacting steam with ethene, C_2H_4

The equation for this reaction is



State the conditions of temperature and pressure used in this process. (2 marks)

temperature _____

pressure _____

- (d) When ethanol is heated with acidified potassium dichromate(VI), it is oxidised to ethanoic acid.

- (i) State the colour change that occurs in the potassium dichromate(VI) during this reaction. (1 mark)

from _____

to _____

(Question continues on next page)

(Turn over)

- (ii) The structural formula of ethanoic acid is CH_3COOH

Draw the displayed formula of ethanoic acid.
(2 marks)

- (iii) Complete the equation for the reaction of ethanoic acid with sodium. (2 marks)



(TOTAL FOR QUESTION 6 = 14 MARKS)

(Questions continue on next page)

(Turn over)

7 Dinitrogen tetraoxide, N_2O_4 , is a colourless gas.

Nitrogen dioxide, NO_2 , is a brown gas.

The two gases can exist together in dynamic equilibrium according to the equation

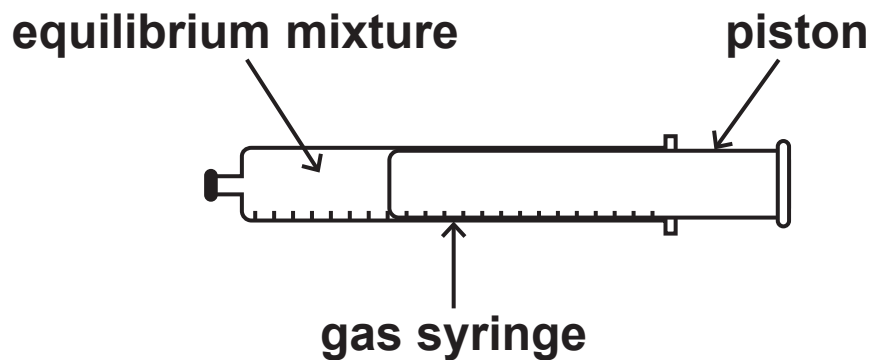


(a) Explain what is meant by the term
DYNAMIC EQUILIBRIUM. (2 marks)

(Question continues on next page)

(Turn over)

- (b) Some N_2O_4 and some NO_2 are put into a sealed gas syringe and allowed to form an equilibrium mixture.



This equilibrium mixture is brown.

(Question continues on next page)

- (i) The pressure of the gas in the syringe is increased by pushing in the piston.
The mixture is then allowed to reach a new equilibrium at the same temperature as before.

Explain why the new equilibrium mixture contains less NO_2 than the original equilibrium mixture. (2 marks)

(Question continues on next page)

(Turn over)

- (ii) A student suggests that the new equilibrium mixture would be lighter in colour than the original equilibrium mixture, as there is now less NO_2 present.

Suggest why the new equilibrium mixture is actually darker than the original. (1 mark)

- (c) Carbon monoxide, CO , and oxides of nitrogen are produced in a car engine when petrol is burned.

These oxides can be partly removed by using a catalytic converter fitted to the car's exhaust system.

(Question continues on next page)

- (i) State how oxides of nitrogen are produced in the car engine. (1 mark)

- (ii) Give a disadvantage of allowing oxides of nitrogen to escape into the atmosphere. (1 mark)

- (iii) Write a chemical equation for the reaction between nitrogen monoxide, NO, and carbon monoxide to form carbon dioxide and nitrogen. (1 mark)

(TOTAL FOR QUESTION 7 = 8 MARKS)

(Questions continue on next page)

(Turn over)

- 8 The concentration of NaClO(aq) in a solution of bleach is found by reacting it with hydrochloric acid.

The equation for the reaction is



An excess of dilute hydrochloric acid is added to 4.00 cm^3 of bleach solution.

60.0 cm^3 of chlorine gas is produced.

- (a) Explain a safety precaution that should be taken when doing this experiment. (2 marks)

(Question continues on next page)

(Turn over)

- (b) (i) Calculate the amount, in moles, of chlorine gas produced. Assume one mole of chlorine gas occupies $24\,000\text{ cm}^3$. (2 marks)

amount of chlorine = _____ mol

- (ii) Determine the amount, in moles, of NaClO in 4.00 cm^3 of bleach. (1 mark)

amount of NaClO = _____ mol

- (iii) Calculate the concentration, in mol/dm^3 , of the bleach solution. (2 marks)

concentration = _____ mol/dm^3

(TOTAL FOR QUESTION 8 = 7 MARKS)

TOTAL FOR PAPER = 70 MARKS
END